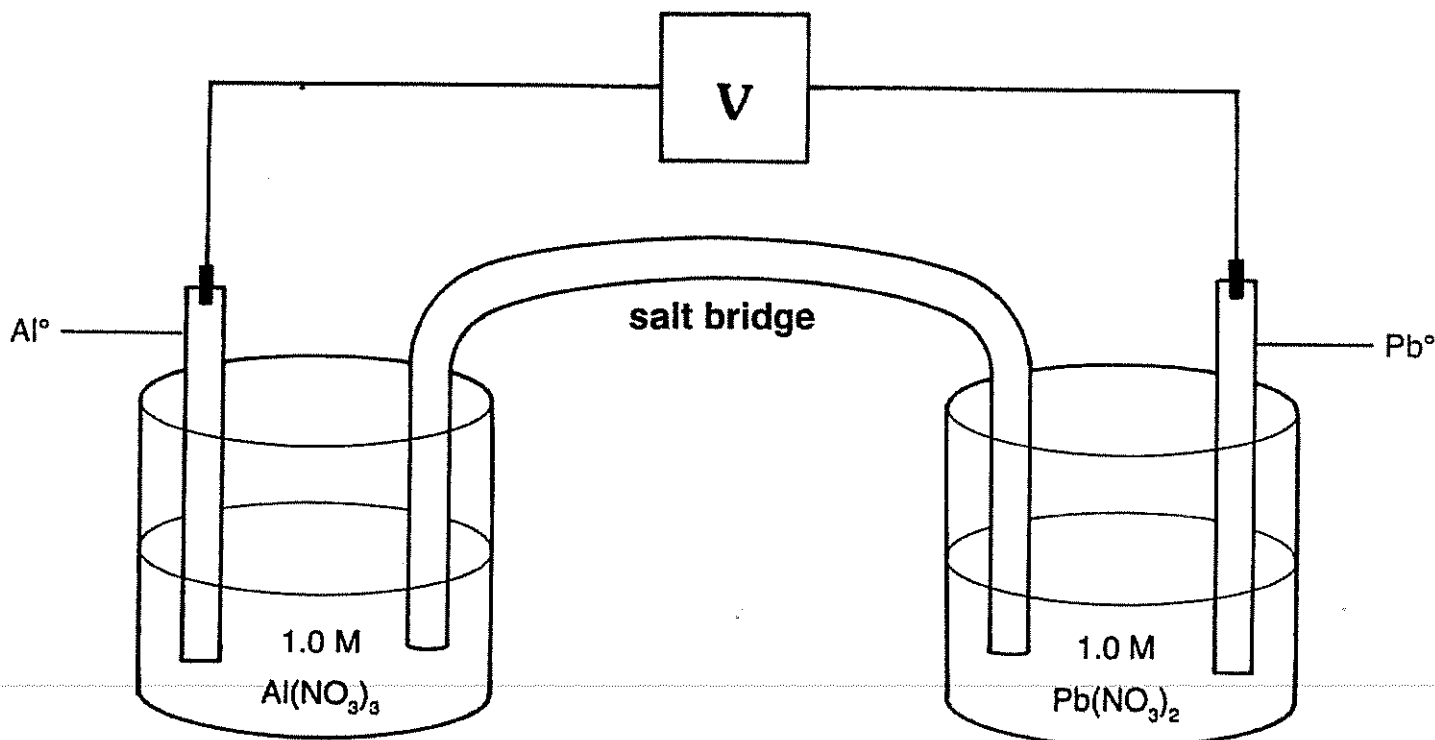


THE ELECTROCHEMICAL CELL

Name _____



Answer the questions below referring to the above diagram and a Table of Standard Electrode Potentials.

1. Which is more easily oxidized, metal, aluminum or lead? _____
2. What is the balanced equation showing the spontaneous reaction that occurs?

3. What is the maximum voltage that the above cell can produce? _____
4. What is the direction of electron flow in the wire? _____
5. What is the direction of positive ion flow in the salt bridge? _____
6. Which electrode is decreasing in size? _____
7. Which electrode is increasing in size? _____
8. What is happening to the concentration of aluminum ions? _____
9. What is happening to the concentration of lead ions? _____
10. What is the voltage in this cell when the reaction reaches equilibrium? _____
11. Which is the anode? _____
12. Which is the cathode? _____
13. What is the positive electrode? _____
14. What is the negative electrode? _____