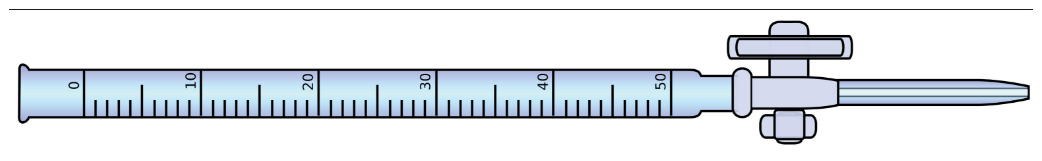
**Titration Quiz Practice Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

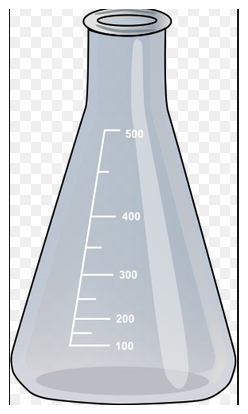
1. A student performs a titration to determine the concentration of a base. Use the information below to calculate the molarity of the base.

Write the balanced equation:

Calculate the M of the base.



The acid is 2.7 M HBr and the amount used in the buret is 50.3 mL.

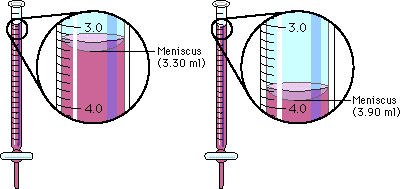


The volume of base, Mg(OH)2, used was 22.7 mL. Indicator was also added.

2. A student performs a titration between KOH and H2SO4. The balanced equation is provided below.

2KOH + H2SO4 🡪 K2(SO4) + 2H2O

The student fills the buret with H2SO4. The starting buret and ending buret volumes are shown below. The student uses a sample of base with a volume of 5.0 mL. What is the concentration, Molarity, of the base?



Start



End